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The Clausius-Clapeyron equation: $dP_i^{sat} = \frac{h_i^{vap}}{RT^2} P_i^{sat} dT$ or $\ln P_i^{sat} = \frac{h_i^{vap}}{R} \left(\frac{1}{T} - \frac{1}{T^*} \right) + \ln P_i^{sat}(T^*)$
101 kPa [] " # \$ % & ' = () h_i^{vap} R 1 T
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thermodynamic equilibrium any
thermodynamic variable for a pure
substance, like pure water, can be
written in terms of any two other
thermodynamic variables, i.e. $p=p(p,T)$
(6.1.1) where the functional relationship
in depends on the substance.

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