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solutions when added to water. (i) Write an equation for the reaction of each acid with water. Hydrogen fluoride, HF, with water: Hydrogen bromide, HBr, with water: Chemistry 3.6 AS 91392 Demonstrate understanding of equilibrium principles in aqueous systems Candidates are expected to recognise common strong acids (HCl, HBr, HNO 3, 2 SO 4 3 +

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aqueous systems WORKBOOK Working to Excellence & NCEA Questions 11. CONTENTS (including NCEA Questions under each topic) 1. Writing Excellence answers to Solubility of sparingly soluble salts questions 2. NCEA Questions for Solubility of sparingly soluble salts 3. Writing Excellence answers to Solubility and Equilibrium questions 4.

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160 Guided Reading and Study Workbook Name ____ Date ____ Class ____ CHAPTER 15,Water and Aqueous Systems(continued) 6. Circle the letter next to each sentence that describes a result of the surface tension of water. a. In a full glass of water, the water surface seems to bulge over the rim of the glass.

SECTION 15.1 WATER AND ITS PROPERTIES (pages 445-449)

Water Aqueous Systems Worksheet Answers Chapter 15 Water And Aqueous Systems Workbook Answers chapter 15 water and aqueous Chapter 15 Water and Aqueous Systems. Chapter 15 “Water and Aqueous Systems”. The Water Molecule: a Review. • Water is a simple tri-atomic molecule, H. 2. O. •Each O-H bond is highly polar, because of the high ...

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File Type PDF Chemistry Workbook Chapter 15 Water And Aqueous Systems Answers Ex. 1: Water acts as base, accepting H + from HCl due to attraction between H + and lone pair of e-in O HCl (aq) + H 2O (l)

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